**Môn hóa:**. GOOD STUDENT PERFORMANCE ISSUES – PART Chemistry:

**I.QUIZZES:**

1- Which of these three oxides. CaO, Fe2O3, SO2, MgO can react with sodium hydroxide?

……………………………………………………………………………………….

2- 200 ml HCl solution of 3,5M is sufficient to dissolve 20 g of the mixture of two oxides CuO and Fe2O3. Calculate the amount of each oxide in the initial mixture?

……………………………………………………………………………………….

Amount of each oxide :

3- The following gases are given: CO2, H2, O2, SO2. Identify the ones with the following properties. Lighter than air.

……………………………………………………………………………………….

4- The following chemicals are given: CuO, Mg, Al2O3, Fe(OH)3, Fe2O3. Choose from the list a substance that reacts with HCl solution to from. A blue solution.

……………………………………………………………………………………….

5- 10,8 grams of metal M (III) reacts with an excess of chlorine and 53,4 grams of salts are obtained. Determine metal M in the reaction.

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6. When experiments completed, residual chlorine gas is eliminated by bubbling chlorine gas into.

……………………………………………………………………………………….

7- The periodic table consists of……………. Periods.

……………………………………………………………………………………….

8- Which chemical is used to engrave letters and images on glass objects?

……………………………………………………………………………………….

9-There are chemical fertilizers: KCl, NH4NO3, NH4Cl, (NH4)2SO4, Ca3(PO4)2, KNO3, Ca(H2PO4)2. Double nutrient fertilizer are?

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10- There is a AlCl3 solution containing CuCl2 impurity. Which chemicals below can be used to purify aluminum salf? Explain and write chemical equations

……………………………………………………………………………………….

**II. SELF COMMENT:**

To 6,3 gram of a mixture X consisting of an alkali metal M and a metal M’ chemotherapy II ( soluble in water) in an excess amount in the of water collccted TOR 3,36 liter H2 and a A solution. Neutral A with excess HCl solution, and then influential solution obtained a gram of dry solids .Find a.

**ĐÁP ÁN:**

**I.QUIZZES:**

**1- SO2**

**2- %mCuO= 20% and %mFe2O3 = 80%**

**3-H2**

**4- CuO**

**5- Al**

**6- NaOH solution**

**7- 7**

**8- HF**

**9- KNO3**

**10- Al**

**II. SELF COMMENT:**

Call in the mixture X:

nM: x mol

nM’: y mol

Call the metal with the same sign as the molar mass=>

According to the article: Mx + M’y = 6,3 g (1\*)

nH2= 3,36/ 22,4= 0,15 (mol)

Chemical equation:

M + 2H2O 🡪 M(OH)2 + H2

x mol x x mol

M’ + 2H2O 🡪 M’(OH)2 + H2

y mol y y mol

According to the chemical equation => nH2 = x+y = 0,15 (mol) (2\*)

Neutral A with excess HCl solution=> Chemical equation:

M(OH)2 + 2HCl🡪 MCl2 +2 H2O

x 2x x mol

M’(OH)2 + 2HCl 🡪 M’Cl2 + 2H2O

y 2y y mol

According to the chemical equation:

Find a.

(M + 71 )x + (M’ + 71) y= a (g)=> Mx + M’y + 71(x + y)

(1\*) and ( 2\* )=> 6,3 + 71x 0,15= 16,95( g ) => a= 16,95 (g)

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Tài liệu được chia sẻ bởi Website VnTeach.Com

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