A. EXPERIMENT

Question 1: Acid oxide is a compound of non-metal elements with an element?

*Answer: ………..*

Question 2: Which of the following gases causes greenhouse effect ?

*Answer: ………..*

Question 3: An oxide of phosphorus with a percentage of P is 43.66%. Know the mass of the oxide is 142 dV. The chemical formula for the oxide is?

*Answer: ………..*

Question 4: There are three white oxides: MgO, Al2O3, Na2O. What reagents can be identified with?

*Answer: ………..*

Question 5: Completely absorbed 11.2 liters of CO2 (tc) with a solution containing 20 g of NaOH. Is salt made up?

*Answer: ………..*

Question 6: What is the product of the decomposition of canxicacarbins by heat?

*Answer: ………..*

Question 7: Dissolve 12.4 grams of Nitric oxide in water to obtain 500 ml of solution A. What is the molar concentration of solution A?

*Answer: ………..*

Question 8: Dissolve 5.6 grams of CaO in 14.6% HCl solution. What is the volume of HCl solution used?

*Answer: ………..*

Question 9: To identify 3 colorless gases: SO2, O2, H2 contained in 3 bottles lost the label we use?

*Answer: ………..*

Question 10: What is the mass of gas for hydrogen by 32?

*Answer: ………..*

B. DISCUSSION

Question 11:

Add 0.83g of the mixture of Al and Fe to the residual solution of excess H2SO4. After the reaction, 0.56 liters of H2 (TOR) were obtained.

a. Write the chemical equations

b. Calculates the percentage by weight of each metal in the mixture

C. ANSWER

|  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Question | 1 | 2 | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10 |
| Answer | Oxi | CO2 | P2O5 | Use water | NaHCO3 | CaO and CO2 | 0.8M | 50 | Dark purple knotted paprika and fire red waxed | SO2 |
| Scores | 5 | 5 | 5 | 5 | 5 | 5 | 5 | 5 | 5 | 5 |

1. Tests

2. self-criticism

|  |  |
| --- | --- |
| Answer | Scores |
| a. equation      2Al + 3H2SO4 (l) → Al2 (SO4) 3 + 3H2 (1)      Fe + 3H2SO4 (l) → FeSO4 + H2 (2) | 10  10 |
| b. nH2 = 0.025 mol Call n Al as a and nFe as b  Pt (1): nH 2 = 3a / 2 (mole)  Pt (2): nH2 = b (mole)  we have pt:           27a + 56b = 0.83           3a / 2 + b = 0.025           a = 0.01           b = 0.015   MAl = 0.27 g   →% Al = 32.53%       % Fe = 67.47% | 10  10  10  10 |

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